CONVENIENT SYNTHETIC ROUTE TO (+) - FRONTALIN

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(\pm)-Frontalin was synthesized in a few steps sequence starting from ethyl acetoacetate involving intramolecular cyclization using palladium chloride as catalyst.

Frontalin, a pheromone of southern pine beetle, <u>Dendroctonus frontalis</u>, was first isolated by Kinzer, ¹⁾ in 1969, and the structure was determined as 1,5-dimethyl-6,8-dioxabicyclo [3.2.1] octane (1) by spectroscopic data and confirmed by a synthesis of its racemate. ^{1,2)} On account of its structural novelty and economic value, this compound has been studied extensively. ³⁻⁵⁾ By now, several syntheses (racemic as well as asymmetric) have reported for frontalin. ⁶⁻¹¹⁾

In the course of our studies on the palladium catalysed cyclisation, we recently reported a three-step synthesis of 1,3-dimethyl-2,9-dioxabicyclo [3.3.1] nonane $(\underline{2})$, an insect attractant. As an extension of this reaction, here, we wish to report a convenient synthetic method for the synthesis of $(\underline{+})$ -frontalin starting from commercially available ethyl acetoacetate.

Alkylation of ethyl acetoacetate ($\underline{3}$) with 4-bromo-1-butene ($\underline{4}$) (NaOEt-EtOH, reflux, 12 h), followed by alkaline hydrolysis and decarboxylation (aq. KOH, then 5% $\mathrm{H_2SO_4}$, reflux, 1 h) gave 6-hepten-2-one ($\underline{5}$) in 71% yield. Ketone $\underline{5}$ was treated with oxysulfurane [(CH $_3$) $_3$ $\underline{\$}$ =0 I $^-$ /NaH] in dry dimethyl sulfoxide at 60 $^{\mathrm{O}}$ C for 1 h gave epoxide $\underline{6}$ in 55% yield. Acid hydrolysis of epoxide (5% $\mathrm{H_2SO_4}$, THF) gave the diol $\underline{7}$ which was cyclised directly to ($\underline{+}$)-frontalin ($\underline{1}$) (50% yield) using palladium chloride as catalyst in anhydrous dimethoxyethane with copper (II) chloride as reoxidant for palladium. The spectral properties (IR, 1 H and 13 C NMR, MS) of the product were identical with those reported for $\underline{1}$. 1 , 6 , 9 , 13)

$$\frac{4}{3}$$
NaOEt/EtOH,
$$\frac{6}{100}$$
NaOEt/EtOH,
$$\frac{1}{100}$$
NaOH/EtOH,
$$\frac{1}{100}$$
NaOH/EtOH

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